

## Interpretation of water analysis for irrigation

### Electrical Conductivity (EC)

Electrical conductivity (EC) is the ability or inability of a substance to allow better flow of electricity. Pure water strongly resists the passage of an electric current. However, when salts are dissolved in water they improve its conductivity, so the greater the quantity of dissolved salts water contains the higher will be its conductivity reading. A measure of electrical conductivity is therefore commonly used as a fairly reliable indicator of the degree of salinity of a water sample. It does not identify the specific dissolved salts, or the affects on crops or soils.

### EC units

The units in which electrical conductivity can be expressed are deciSiemens per metre (dS/m) although it can also be expressed in microSiemens per centimetre ( $\mu\text{S/cm}$ ) or milliSiemens per centimetre (mS/cm). To add to the confusion electrical conductivity is sometimes also expressed as milli mho per cm. The electrical conductivity reading is temperature dependant and EC thresholds are usually based on a temperature based on 25 °C.

**Table 1:** When checking water analysis data make sure you know what the units are:

$\mu\text{S/cm}$ = microSiemen per centimetre
mS/cm = milliSiemen per centimetre
dS/cm = deciSiemen per centimetre
dS/m = deciSiemen per meter
mmhos/cm = milli mhos per centimetre

**Table 2:** To convert between the various units:

1 mS/cm = 1000 $\mu\text{S/cm}$
1 mS/cm = 1 dS/m
1 dS/m = 1 mmhos/cm

**Table 3:** Some example electrical conductivity values of various water sources

Water type	Electrical Conductivity (EC)	
	( $\mu\text{S/cm}$ )	(dS/m)
Distilled water	1	0.001
Pure rainwater	30	0.03
Sewage effluent	840	0.84
Great Artesian Basin water	700-1000	0.7-1.0
Freshwater	0-1500	0-1.5
Brackish water	1500-15000	1.5-15
Upper limit for drinking water	1600	1.6
Saline water	>4800	4.8
Seawater	55000	55
Industrial waters	100-10000	0.1-10

## Total Dissolved Salts (TDS)

Total Dissolved Salts is the sum of all the ions present in a sample of water and represents the total salt content of the water. The total salt concentration of the tested water is one of the most important pieces of information presented in the water analysis report. High levels of soluble salts can induce physiological drought in the plant. Turf roots may have an adequate water supply, but are unable to absorb the water due to osmotic pressure.

The units of TDS are parts per million or milligrams per litre (ppm or mg/L),

**Table 4:** The units of TDS are:

mg/L – milligram per litre
ppm – parts per million
1 mg/L = 1 ppm

Table 5 lists the levels at which the salt concentration poses a hazard when considering the water for an irrigation source.

**Table 5:** Salinity levels posing a hazard in irrigation water.

TDS ppm or mg/L	EC dS/m or mmhos/cm	Salinity Hazard
<500	<0.8	Low
500-1000	0.8-1.6	Medium
1000-2000	1.6-3	High
>2000	>3	Very High

## Conversion between TDS and EC

The total salt concentration can either be expressed as total dissolved salts (TDS) or Electrical Conductivity (EC). Both measures may be presented on the report.

**Table 6:** Conversion from TDS to EC

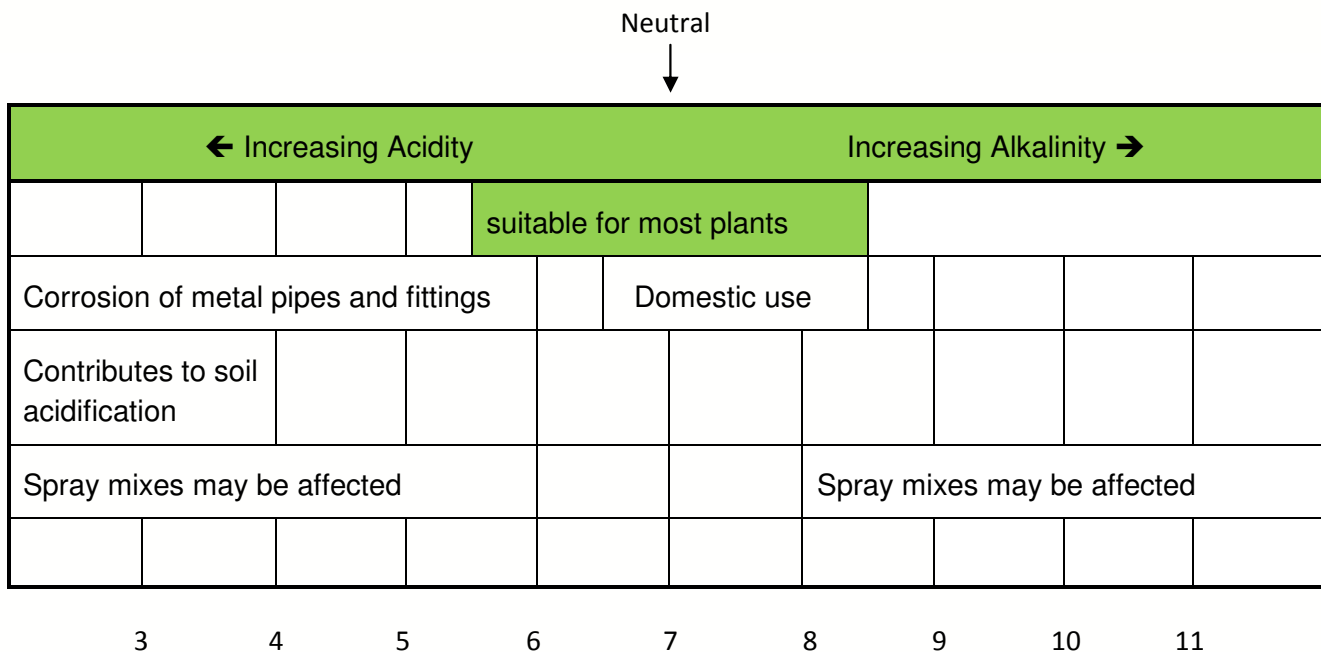
$TDS \text{ in ppm or mg/L} \approx 640 \times EC_w \text{ in dS/m or mmhos/cm}$
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## pH

The term pH is a measure of the acidity or alkalinity of a water sample. The pH of natural waters normally falls between the range of pH 4.0 to pH 9.0. Soils are generally highly buffered systems and the pH of the soil would not be significantly affected by the application of irrigation water within this range. Waters having pH values greater than 8.0 would be expected to contain carbonates and bicarbonates, the calcium form of which can precipitate and block equipment. The waters usefulness would depend on the relative amounts of these salts. Corrosion is more rapid in acid than in neutral or alkaline waters. Irrigation with strongly acid water may dissolve iron, aluminum and magnesium from the soil in amounts that could be toxic to plant growth.

The test for pH measures the balance between positive hydrogen ions ( $H^+$ ) and negative hydroxyl ions ( $OH^-$ ). This indicates if water is alkaline ( $pH > 7$ ), neutral (7) or acid ( $< 7$ ) as shown in figure 1.

**Figure 1 - Degree of Acidity or Alkalinity**



### Turbidity

Turbidity is a measure of water clarity, and is an indicator of how much solid matter (such as clay, silt, organic matter and micro-organisms) is suspended in the water.

### Hardness

There are several definitions of water hardness – it could mean the water has high concentrations of ion, manganese, sulphates, carbonates and/or bicarbonates. Or it could specifically refer to the concentration of calcium or calcium carbonates. Without knowing what this refers to it is difficult to identify, below are some simple limits that may be of use.

Bicarbonate concentrations between 90 and 200 mg/L can cause increasing plant growth problems and cause foliage or container staining, > 500mg/L is unsuitable for nursery irrigation. Calcium is required for plant growth in low concentrations and is not considered toxic but high concentrations can affect the calcium:magnesium ratio and cause scale build up. Maintaining pH below 7.2 will prevent scale formation. Calcium carbonate (Alkalinity) levels greater than 125 mg/L may cause pH to rise to unacceptable levels. Above 500 mg/L will cause severe problems and is not suitable. High concentrations will cause scale build up in irrigation systems.

### Specific ions

Specific ions can be toxic to plants and/or detrimental to the soil physical structure. Certain ions (Na, Cl, and B) can cause direct root injury, accumulate in shoot tissues and cause shoot toxicity problems, or cause direct foliar toxicity on plant leaves. These problems are almost always present when high total salinity is present.

# FACT SHEET

NO: 013

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**Turf**  
**Queensland**  
Grow Green

*Soluble ions found in water are:*

## **Cations**

Calcium (Ca<sup>2+</sup>)

Magnesium (Mg<sup>2+</sup>)

Sodium (Na)

Potassium (K)

## **Anions**

Carbonates (CO<sub>3</sub><sup>2-</sup>)

Bicarbonates (HCO<sub>3</sub><sup>1-</sup>)

Chloride (Cl<sup>1-</sup>)

Sulfate (SO<sub>4</sub><sup>2-</sup>)

Nitrate (NO<sub>3</sub><sup>1-</sup>)

Boron (B<sup>3+</sup>)

Phosphate (PO<sub>4</sub><sup>3-</sup>)

## **Iron (Fe<sup>3+</sup>)**

Dissolved iron in water is present in the ferrous state. Except at low pH values, ferrous iron is readily oxidised to ferric iron, an insoluble reddish brown precipitate on exposure to air and sunlight. For information on how to take water samples for iron analysis, check with the laboratory that will be conducting the analysis.

For localised irrigation schemes that use very small emitters, the presence of iron in the irrigation water has proved to be a problem. It has been found that iron bacteria flourish in water that contains as little as 1.0 mg/L of iron. The bacteria extract the iron out of solution and convert it into a rust coloured sludge, which quickly blocks filters and emitters.

## **Nitrate (NO<sub>3</sub><sup>-</sup>)**

For many crops nitrate in the irrigation water will provide them with some extra nitrogen. Nitrate sensitive crops could be affected by concentrations greater than 22 mg/L nitrate and problems may occur with increasing concentrations up to 133 mg/L nitrate, above which severe problems could arise.

Nitrogen could be found in dams containing decaying organic matter; or in underground water contaminated with seepage from soils, that have had large quantities of nitrogen fertiliser applied, or by effluent from various sources.

## **Sodium (Na<sup>+</sup>)**

Symptoms of sodium toxicity appear as burning or drying on the outer edges of older leaves, progressing inwards towards the centre. Sodium can be absorbed through roots or leaves if sprinkler irrigation is used. Tree crops and woody perennials are most affected. With the exception of beans annual crops are not so sensitive.

## **Manganese**

Dissolved manganese can precipitate out to block irrigation equipment and cause black bacterial slime to grow within the irrigation system reducing efficiency.

- 0.05 mg/L can lead to slime build up in irrigation systems and is the maximum level for disposal into streams and waterways.
- 0.2 mg/L is the maximum concentration for irrigation. Higher levels are toxic to plants. 1.5 mg/L will clog irrigation equipment.

